

<p>How many moles of sodium atoms correspond to 1.56×10^{21} atoms of sodium?</p> <p>$2.59 \times 10^{-3} \text{ mol Na}$</p>	<p>How many moles of Al are in 2.16 mol of Al_2O_3?</p> <p>4.32 mol Al</p>	<p>How many moles of H_2 and N_2 can be formed by the decomposition of 0.145 mol of ammonia, NH_3?</p> <p>0.0725 mol N_2, 0.218 mol H_2</p>	<p>What is the total number of atoms in 0.260 mol of glucose, $\text{C}_6\text{H}_{12}\text{O}_6$?</p> <p>$1.57 \times 10^{23}$ atoms</p>
<p>What is the mass of 1.00 mol of Sodium?</p> <p>23.0g Na</p>	<p>What is the mass of 1.00 mol of Sulfur?</p> <p>32.1g S</p>	<p>What is the mass of 1.00 mol of Chlorine?</p> <p>35.5g Cl</p>	<p>Determine the mass in grams of 1.35 mol Fe</p> <p>75.4g Fe</p>
<p>Determine the mass in grams of 24.5 mol O</p> <p>392g</p>	<p>Determine the mass in grams of 0.876 mol Ca</p> <p>35.1g Ca</p>	<p>Determine the mass in grams of 1.25 mol $\text{Ca}_3(\text{PO}_4)_2$</p> <p>388g $\text{Ca}_3(\text{PO}_4)_2$</p>	<p>Determine the mass in grams of 0.625 mol $\text{Fe}(\text{NO}_3)_3$</p> <p>151g $\text{Fe}(\text{NO}_3)_2$</p>
<p>Determine the mass in grams of 0.600 mol C_4H_{10}</p> <p>34.9g C_4H_{10}</p>	<p>Determine the mass in grams of 1.45 mol $(\text{NH}_4)_2\text{CO}_3$</p> <p>139g $(\text{NH}_4)_2\text{CO}_3$</p>	<p>Calculate the number of moles of 21.5 g CaCO_3</p> <p>0.215 mol</p>	<p>Calculate the number of moles of 1.56 g NH_3</p> <p>0.0916 mol</p>
<p>Calculate the number of moles of 16.8 g $\text{Sr}(\text{NO}_3)_2$</p> <p>0.0794 mol</p>	<p>Balance the following reaction:</p> $\text{Ca}(\text{OH})_2 + \text{HCl} \rightarrow \text{CaCl}_2 + \text{H}_2\text{O}$ $\text{Ca}(\text{OH})_2 + 2\text{HCl} \rightarrow \text{CaCl}_2 + 2\text{H}_2\text{O}$	<p>Balance the following reaction:</p> $\text{AgNO}_3 + \text{CaCl}_2 \rightarrow \text{Ca}(\text{NO}_3)_2 + \text{AgCl}$ $2\text{AgNO}_3 + \text{CaCl}_2 \rightarrow \text{Ca}(\text{NO}_3)_2 + 2\text{AgCl}$	<p>Balance the following reaction:</p> $\text{Fe}_2\text{O}_3 + \text{C} \rightarrow \text{Fe} + \text{CO}_3$ $\text{Fe}_2\text{O}_3 + \text{C} \rightarrow 2\text{Fe} + \text{CO}_3$

<p>Write and balance the following: Sodium Bicarbonate and Sulfuric Acid react to produce Sodium Sulfate, Water, and Carbon Dioxide</p> $2\text{NaHCO}_3 + \text{H}_2\text{SO}_4 \rightarrow \text{Na}_2\text{SO}_4 + 2\text{H}_2\text{O} + 2\text{CO}_2$	<p>Write and balance the following: Aluminum Oxide and Sulfuric Acid react to produce Aluminum Sulfate and Water</p> $\text{Al}_2\text{O}_3 + 3\text{H}_2\text{SO}_4 \rightarrow \text{Al}_2(\text{SO}_4)_3 + 3\text{H}_2\text{O}$	<p>Write and balance the following: Sodium Carbonate and Water react to produce Carbonic Acid and Sodium Oxide</p> $\text{Na}_2\text{CO}_3 + \text{H}_2\text{O} \rightarrow \text{H}_2\text{CO}_3 + \text{Na}_2\text{O}$	<p>Write and balance the following: Magnesium Phosphate and Potassium Sulfide react to produce Magnesium Sulfide and Potassium Phosphate</p> $\text{Mg}_3(\text{PO}_4)_2 + 3\text{K}_2\text{S} \rightarrow 3\text{MgS} + 2\text{K}_3\text{PO}_4$
<p>$\text{Na}_2\text{S}_2\text{O}_3 + 4\text{Cl}_2 + 5\text{H}_2\text{O} \rightarrow \text{NaHSO}_4 + 8\text{HCl}$</p> <p>How many moles of $\text{Na}_2\text{S}_2\text{O}_3$ are needed to react with 0.12mol of Cl_2?</p> <p>0.030mol $\text{Na}_2\text{S}_2\text{O}_3$</p>	<p>$\text{Na}_2\text{S}_2\text{O}_3 + 4\text{Cl}_2 + 5\text{H}_2\text{O} \rightarrow \text{NaHSO}_4 + 8\text{HCl}$</p> <p>How many moles of $\text{Na}_2\text{S}_2\text{O}_3$ are needed to react with 0.12mol of Cl_2?</p> <p>0.24mol HCl</p>	<p>$\text{Na}_2\text{S}_2\text{O}_3 + 4\text{Cl}_2 + 5\text{H}_2\text{O} \rightarrow \text{NaHSO}_4 + 8\text{HCl}$</p> <p>How many moles of H_2O are required for the reaction of 0.12mol of Cl_2?</p> <p>0.15mol H_2O</p>	<p>$\text{Na}_2\text{S}_2\text{O}_3 + 4\text{Cl}_2 + 5\text{H}_2\text{O} \rightarrow \text{NaHSO}_4 + 8\text{HCl}$</p> <p>How many moles of H_2O react if 0.24mol HCl is formed?</p> <p>0.15mol H_2O</p>
<p>$3\text{Cu} + 8\text{HNO}_3 \rightarrow 3\text{Cu}(\text{NO}_3)_2 + 2\text{NO} + 4\text{H}_2\text{O}$</p> <p>How many grams of HNO_3 are needed to dissolve 11.45g of Cu?</p> <p>30.31g HNO_3</p>	<p>$2\text{Al} + \text{Fe}_2\text{O}_3 \rightarrow \text{Al}_2\text{O}_3 + 2\text{Fe}$</p> <p>Calculate the mass of iron that can be formed with 1.75 mol Fe_2O_3</p> <p>195g Fe</p>	<p>How many molecules of water are produced when 8.92 liters of oxygen are reacted with excess hydrogen?</p> <p>4.79×10^{23} molecules H_2O</p>	<p>When 47.6 g of calcium chlorate decomposes into calcium chloride and oxygen, how many liters of oxygen are produced?</p> <p>15.5L O_2</p>
<p>How many grams of aluminum chloride are produced when 11.82 grams of aluminum are reacted with excess chlorine gas?</p> <p>58.56g AlCl_3</p>	<p>Magnesium sulfate reacts with sodium chloride to produce magnesium chloride and sodium sulfate. How many grams of magnesium chloride are produced if 5.75 grams of sodium chloride is used?</p> <p>4.68g MgCl_2</p>	<p>How many grams of iron(III) sulfide are produced when 125 grams of iron react with sulfur?</p> <p>233 g Fe_2S_3</p>	<p>How many molecules of Barium hydroxide are needed to produce 3.33 grams of water?</p> $\text{Ba}(\text{OH})_2 + \text{HNO}_3 \rightarrow \text{H}_2\text{O} + \text{Ba}(\text{NO}_3)_2$ <p>5.56×10^{22} molecules $\text{Ba}(\text{OH})_2$</p>