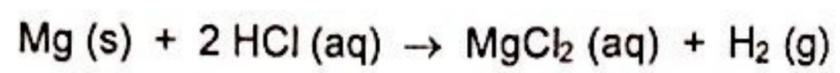


Name KEY

Period _____

STOICHIOMETRY WORKSHEET (MOLE-MOLE)

1. Magnesium reacts with hydrochloric acid according to the following balanced chemical equation:



If two moles of hydrochloric acid react with excess magnesium, how many moles of hydrogen gas will be produced?

$$\frac{2 \text{ mole HCl}}{2 \text{ mole HCl}} \left| \frac{1 \text{ mole H}_2}{1 \text{ mole H}_2} \right. = \boxed{1 \text{ mole H}_2}$$

2. Aluminum reacts with HCl to produce aluminum chloride and hydrogen gas. Write a balanced equation for the reaction and calculate the number of moles of HCl required to react with 0.87 moles of Al.



* DO NOT
REDUCE YOUR
RATIOS!! *

$$\frac{0.87 \text{ moles Al}}{2 \text{ mole Al}} \left| \frac{6 \text{ mole HCl}}{1 \text{ mole Al}} \right. = \boxed{2.61 \text{ mole HCl}}$$

3. Glucose (C₆H₁₂O₆) combines with O₂ in the body to produce carbon dioxide and water. Write a balanced equation for this reaction. How many moles of O₂ are required to combine with 0.25 moles of glucose? How many moles of CO₂ and H₂O would be produced in this reaction?



$$\frac{0.25 \text{ mole C}_6\text{H}_{12}\text{O}_6}{1 \text{ mole C}_6\text{H}_{12}\text{O}_6} \left| \frac{6 \text{ mole O}_2}{1 \text{ mole C}_6\text{H}_{12}\text{O}_6} \right. \Rightarrow \boxed{1.5 \text{ mole O}_2}$$

* Same answer for CO₂ & H₂O b/c same ratio.

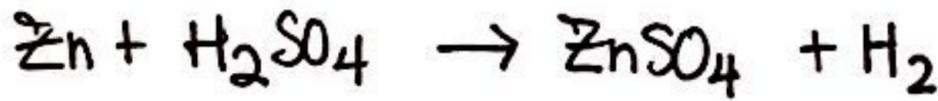
4. Calcium carbonate combines with HCl to produce calcium chloride, water, and carbon dioxide gas. Write the balanced equation for this reaction. How many moles of HCl are required to react with 2.5 moles of calcium carbonate? How many moles of carbon dioxide would be produced?



$$\frac{2.5 \text{ moles CaCO}_3}{1 \text{ mole CaCO}_3} \left| \frac{2 \text{ mole HCl}}{1 \text{ mole CaCO}_3} \right. = \boxed{5 \text{ mole HCl}}$$

$$\frac{2.5 \text{ mole CaCO}_3}{1 \text{ mole CaCO}_3} \left| \frac{1 \text{ mole CO}_2}{1 \text{ mole CaCO}_3} \right. = \boxed{2.5 \text{ mole CO}_2}$$

5. Zinc reacts with sulfuric acid (H₂SO₄) to yield zinc sulfate and hydrogen gas. How many moles of hydrogen will be produced if 0.36 moles of zinc react with an equal amount of H₂SO₄?



$$\frac{0.36 \text{ moles Zn}}{1 \text{ mole Zn}} \left| \frac{1 \text{ mole H}_2}{1 \text{ mole Zn}} \right. = \boxed{0.36 \text{ mole H}_2}$$